Arrhenius equation with a base temperature of 30°C (Safley and Westerman 1990):

$$f = \exp\left[\frac{E(T_2 - T_1)}{RT_1T_2}\right] \qquad (4)$$

Where

f	=	van't Hoff-Arrhenius factor (unitless)
T_1	=	303.15K
T_2	=	Ambient temperature (K)
E	=	Activation energy constant (15,175 cal/mol)
R	=	Ideal gas constant (1.987 cal/K mol)